Common Polyatomic Ions

+1	-1	-2	-3
NH_4^+ (ammonium)	OH^- (hydroxide)	CO_3^{2-} (carbonate)	PO_4^{3-} (phosphate)
	NO_3^- (nitrate)	SO_4^{2-} (sulfate)	
	ClO_3^- (chlorate)	CrO_4^{2-} (chromate)	
	ClO_4^- (perchlorate)	$Cr_2O_7^{2-}$ (dichromate)	
	$C_2H_3O_2^-$ (acetate)		
	HCO_3^- (hydrogen carbonate)		

Diatomic Gases: H_2, O_2, F_2, Br_2, I_2, N_2, Cl_2

Capacity of electron shells: 2, 8, 8, 18, 18, 32, 32

Table of Solubilities in Water*

i - insoluble a - soluble	acetate	bromide	carbonate	chloride	chromate	hydroxide	nitrate	phosphate	sulfate	sulfide
Aluminum	i	a		a		i	a	i	a	
Ammonium	a	a	a	a	a	a	a	a	a	a
Barium	a	a	i	a	i	a	a	i	i	
Calcium	a	a	i	a	a	a	a	i	i	
Copper II	a	a	i	a	i	i	a	i	a	i
Iron II	a	a	i	a		i	a	i	a	i
Iron III	a	a		a	i	i	a	i	i	
Lead	a	i	i	i	i	i	a	i	i	i
Magnesium	a	a	i	a	a	i	a	i	a	
Mercury I	i	i	i	i	i		a	i	i	i
Mercury II	a	i	i	a	i	i	a	i		i
Potassium	a	a	a	a	a	a	a	a	a	a
Silver	i	i	i	i	i		a	i	i	i
Sodium	a	a	a	a	a	a	a	a	a	a
Zinc	a	a	i	a	a	i	a	i	a	i

*This table shows which ionic compounds dissolve (or stay dissolved) when in water