Acid Dissociation Constants, K_a at 25ºC

(calculated from pK_a values in Handbook of Chemistry & Physics, CRC, 85th ed.)

Name	Formula	K_{a1}	K_{a2}	K_{a3}
Acetic	$HC_2H_3O_2$	1.8×10^{-5}		
Ascorbic	$HC_6H_7O_6$	9.1×10^{-5}		
Benzoic	$HC_7H_5O_2$	6.5×10^{-5}		
Carbonic	H_2CO_3	4.3×10^{-7}	5.6×10^{-11}	
Chloroacetic	$HC_2H_2O_2Cl$	1.3×10^{-3}		
Citric	$H_3C_6H_5O_7$	7.4×10^{-4}	1.7×10^{-5}	4.0×10^{-7}
Cyanic	HCNO	7.4×10^{-4}		
Formic	$HCHO_2$	1.8×10^{-4}		
Hydrocyanic	HCN	4.9×10^{-10}		
Hydrofluoric	HF	6.8×10^{-4}		
Hydrosulfuric	HS	8.9×10^{-8}		
Hypochlorous	HClO	3.0×10^{-8}		
Hypoiodous	HIO	3×10^{-11}		
Lactic	$HC_3H_5O_3$	1.4×10^{-4}		
Nitrous	HNO_2	4.5×10^{-4}		
Phosphoric	H_3PO_4	6.9×10^{-3}	6.2×10^{-8}	4.8×10^{-13}
Propionic	$HC_3H_5O_2$	1.3×10^{-5}		
Sulfuric	H_2SO_4	Strong	1.0×10^{-2}	

Base Dissociation Constants, K_b at 25ºC

(calculated from pK_a values in Handbook of Chemistry & Physics, CRC, 85th ed.)

Name	Formula	K_b
Ammonia	NH_3	1.8×10^{-5}
Aniline	$C_6H_5NH_2$	7.7×10^{-10}
Ethylamine	$C_2H_5NH_2$	4.5×10^{-4}
Methylamine	CH_3NH_2	4.5×10^{-4}